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Synthesis, structural characterization and biological activity of transition metal complexes of schiff base ligand-furan derivatives

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ABSTRACT

A series of Ti (III), Mn (III), V (III), Fe (III), Co (III), Ru (III), Ru (II), Vo (IV) and Cu (II) complexes have been synthesized from the Schiff base ligand. The Schiff base ligand is derived from 2-furancarboxaldehyde and O-phenylenediamine. The resulting complexes were characterized by elemental analyses, magnetic moment measurement, molar conductance, thermal analyses (TGA), IR, ^{1}H-NMR and solid reflectance. The ligand dissociation, as well as metal ligand stability constants were calculated, pH-metrically at 25^0C and ionic strength $\mu = 0.1$ (IM Nacl). IR spectra show that ligand is coordinated to the metal ions in a tetradenatate manner, with ONNO donor sites of azomethine N and furan O. The molar conductance data reveals that all metal chelates are electrolytes. The electronic spectral data of the complexes displayed the proper transitions and octahedral geometry. The synthesized ligands in comparison to their metal complexes, were also screened for their antibacterial activity against bacterial species, Escherichia coli, pseudomonas aeruginosa and staphyalococcus pyogones. The activity data show the metal complexes to be more potent antibacterials than Schiff base ligand against one or more bacterial species.

Key words: Transition metal complexes, 2-furan castoxaldehyde, O-phenylenediamine, Biological activity.

INTRODUCTION

The chemistry of the carbon-nitrogen double bond plays a vital role in the progress of chemical science[1]. Schiff base compounds have been used as five chemicals and medical substrates. In the field of coordination chemistry, Schiff base metal complexes have a curious history[2,3]. Metal ion play vital roles in the vast number of biological processes. Metal complexes with Schiff base ligand have been studied for their application in biological, clinical, analytical and pharmacological areas[4]. The importance of furan derivatives is quite evident from a number of papers. patents etc. published every year. Furan derivatives exhibiting anti-inflammatory,[5] antituberculosis,[6] anticancer,[7] antimicrobial,[8] antifungal,[9] activity have been reported in literature. Many potent antibacterial and antifungal compounds synthesized by condensation of aldehydes with various heterocyclic have been reported[10-12].

Maganese, Vanadium and Ruthenium Possess as number of oxidation states and have excellent complexing property. The last two metals and their complexes exhibits biological properties[13-14]. Keeping in view these facts we have synthesized ligand having oxygen, nitrogen, donor atoms and study their complexation behaviour and biological activity.

EXPERIMENTAL SECTION

Material and Reagents

All chemicals used were of analytical reagent grade (AR) and of the highest purity available. They included 2-furan carboxaldehyde (sigma), O-phenylenediamine (Aldrich), Titanium (III), chloride (merck p.a.), manganese (III) , chloride (Aldrich p.a.), Vanadium (III) chloride (Aldrich p.a.), Iron (III) chloride (B.D.H.), Cobalt (III) chloride

(Merck p.a.). The organic solvents used included absolute ethyl alcohol, diethylether and dimethylformamide (DMF). These solvents were either spectroscopically pure from BDH or purified by the recommended method[15] and tested for their spectral purity. De-ionized water collected from all-glass equipments was normally used in all preparations. Fresh stock solutions of $1x10^{-3}$M ligand, were prepared by dissolving the accurately weighed amount of Schiff base ligand (0.264 gm/L) in the appropriate volume of absolute ethanol. The $1x10^{-3}$ M stock solutions of metal salts Fe (III), Ti (III), V (III), Mn (III), Co (III), Ru (III), Cu (II) were prepared by dissolving accurately weighed amounts of the metal salts in appropriate volume of de-ionized water.

Instrumentation

Melting points were determined on a JSGW apparatus and are uncorrected. I.R. spectra were recorded using a Perkin Elmer 1600 FT spectrometer ^{1}H-NMR spectra were measured on a Brucker WH-500 MHz spectrometer at Ca 5-15% solution in DMSO – d^6 (T.M.S. as internal standard). Elemental analyses was carried out on vitro EL III elementor. Thin layer chromatography (TLC) was performed on Silica gel G for TLC (merck) and spots were visualized by Iodine vapours.

The molar conductance of solid complexes in DMF was measured using OK – 102 (Hungary) conductivity-meter. Magnetic susceptibility was measured with a faraday balance at room temperature. pH measurements were performed using a metrohm 716 DMS titrino connected to metrohm 728 stirrer. C, H, and N were analyzed using M.L.W. micro-elementary CHN analyses. Thermogravimetric analyses was performed with a DU-Pont 2000 thermal analyses.

Synthesis of ligand

A hot solution (60^0C) of o-phenylenediamine (1.08 g, 10 mmol) was mixed with a hot solution (60^0C) of 2-furancarboxaldehyde (1.92 g, 20 mmol) in 50 mL of ethanol. The resulting mixture was left under reflux for 2 h and the solid product formed was separated by filtration, purified by crystallization from ethanol, washed with diethyl ether and then dried in a vacuum over anhydrous calcium chloride. The yellow product is produced in 80% yield[16].

Synthesis of metal complexes

The metal complexes of the Schiff bases, ligand were prepared by the addition of a hot solution (60^0C) of the appropriate metal chloride, (1 mmol) in an ethanol-water mixture (1:1, 25 mL) to the hot solution (60^0C) of the Schiff bases (0.264 g 2mmol) in the same solvent (25 mL). The resulting mixture was stirred under reflux for 1 h whereupon the complexes precipitated. They were collected by filtration, washed with a 1:1 ethanol water mixture and diethyl ether. The analytical data for C, H, N, and S were repeated twice.

Biological activity

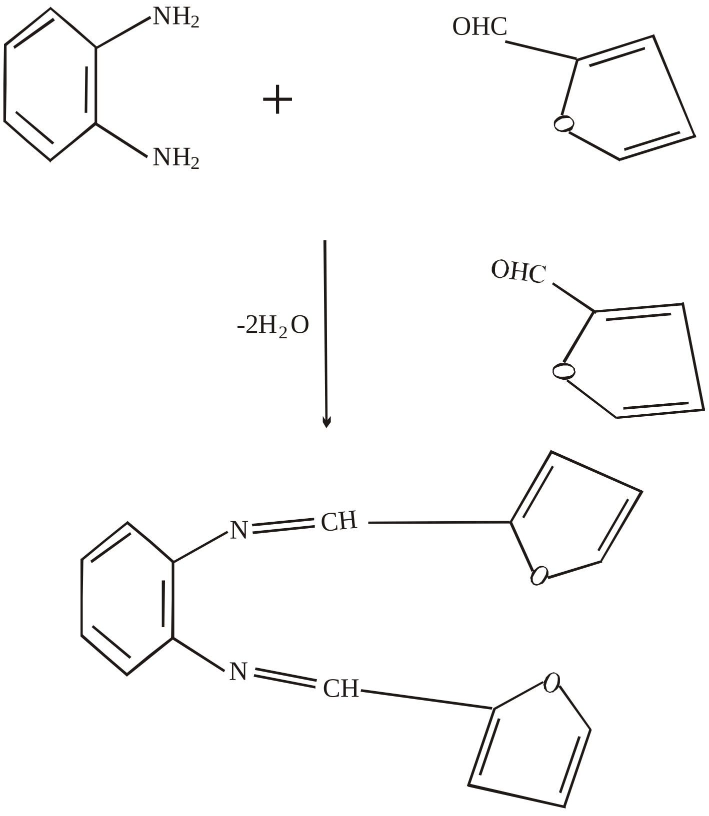
A 0.5 mL spore suspension (10^{-6} – 10^{-7} spore/mL) of each of the investigated organisms was added to a sterile agar medium just before solidification, then poured into sterile Petri dishes (9 cm in diameter) and left to solidify. Using a sterile cork borer (6 mm in diameter), 3 holes (wells) were made in each dish and then 0.1 mL of the tested compounds, dissolved in DMF (100 μg/mL), was poured into these holes. Finally, the dishes were incubated at 37^0C for 48 h. Then clear or inhibition zones were detected around each hole. DMF alone (0.1 mL) was used as a control under the same condition for each organism, and by substrating the diameter of inhibition zone resulting with DMF from that obtained in each case, both antibacterial activities were calculated as a mean of 3 replicates[17-18].

RESULTS AND DISCUSSION

The ligand and its transition metal complexes with Ti (III), V (III), Mn (III), Co (III), VO (IV), Fe (III), Ru (III), Cu (II) & Ru (II) were subjected to elemental analyses where as metal and complexes were estimated gravimetrically in the lab, validated by jobs method. All this analytical data suggested 1:1 M:L stoichiometry for all the complexes.

The M.P. of the ligand and its metal complexes were determined and compared in order to find out the possibilities of formation of complexes. The M.Pt. are given in Table (1). The observed value of magnetic susceptibility was used to calculate to magnetic moment of the complexes. These values suggested paramagnetic nature for Ti (III), Mn (III), V (III), Cr (III), Fe (III) complexes as expected for octahedral d^1, d^4, d^2 d^3 and d^5 complexes. The Co (III) complex is diamagnetic in nature as expected for low spin d^6 ion. The value of magnetic moments of complexes are given in Table (1).

Ti (III), Mn (III), V (III), Ru (III), Fe (III) complexes as expected for octahedral d^1, d^4, d^2 and d^5 complexes. The Co (III) complex is diamagnetic in nature as expected for low spin d^6 ion. The value of magnetic moments of complexes are given in Table (1).



Schiff base Ligand - 2-furancarboxaldehyde-o-phenylene diamine

Figure-1

Table-1: Characterization of Ligand and its Metal Complexes Prepared

S. No.	**Formula of the Ligand and Complex and Molecular Weight**	**Colour**	**M.P. ^{0}C**	**Elemental analyses**					**Molar Conductance ohm^{-1} cm^2 mole^{-1}**	**Magnetic Moments in (B.M.)**
				% of C	**% of H**	**%of N**	**% of M**	**% of Cl**		
1	$C_{16}H_{12}N_2O_2$ Mol. Wt. = 264	Yellow	122	72.73 (72.70)	4.54 (4.45)	10.61 (10.50)	-	-	-	-
2	$[Ti\ (C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 454.5	Yellow	197	42.24 (42.20)	3.52 (3.48)	6.16 (6.12)	10.56 (10.52)	23.43 (23.40)	125	1.77
3	$[Mn(C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 461.5	Dark Brown	202	41.60 (41.56)	3.47 (3.41)	6.07 (6.02)	11.92 (11.88)	23.08 (23.03)	135	5.40
4	$[V(C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 457.5	Light Yellow	199	41.97 (41.90)	3.50 (3.46)	6.12 (6.09)	11.15 (11.10)	23.28 (23.23)	130	2.90
5	$[Fe(C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 462.5	Brown	208	41.51 (41.46)	3.46 (3.40)	6.05 (6.02)	12.11 (12.07)	23.03 (23.00)	140	5.90
6	$[Co(C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 465.5	Dark Brown	212	41.24 (41.20)	3.44 (3.40)	6.01 (6.00)	12.67 (12.61)	22.88 (22.82)	138	Diamagnetic
7	$[Ru(C_{16}H_{12}N_2O_2)(H_2O)_2]\ Cl_3$ Mol. Wt. = 507.5	Green	245	37.83 (37.81)	3.15 (3.11)	5.52 (5.50)	19.90 (19.86)	20.98 (20.94)	142	1.80
8	$[Ru(C_{16}H_{12}N_2O_2)(H_2O)_2]Cl_2$ Mol. Wt. = 472.0	Dull White	239	40.68 (40.61)	3.39 (3.35)	5.93 (5.90)	21.40 (21.38)	15.04 (15.01)	80	Diamagnetic
9	$[VO(C_{16}H_{12}N_2O_2)(H_2O)]Cl_2$ Mol. Wt. = 420.0	Yellowish Orange	190	45.71 (45.62)	3.33 (3.25)	6.66 (6.60)	12.14 (12.10)	16.90 (16.82)	85	1.71
10	$[Cu(C_{16}H_{12}N_2O_2)(H_2O)_2]Cl_2$ Mol. Wt. = 434.5	Greenish Yellow	195	44.18 (44.11)	3.68 (3.62)	6.44 (6.38)	14.61 (14.56)	16.34 (16.29)	90	1.88

Figures in parenthesis are observed values.

By using the relation λ_m= K/C, the molar conductance of the complexes (λ_m) can be calculated, where C is the molar concentration of the metal complex solutions. The chelates were dissolved in DMF and the molar conductivities of 10^{-3} M of their solutions at 25± 2 ^{0}C were measured. Tables-1 show the molar conductance values of the complexes. The value of molar conductance indication the 1:3 and 1:2 electrolytic nature of synthesized complexes.

Electronic spectra

The electronic spectrum of the complex of Ti (III) exhibits a single broad band at 19250 Cm^{-1} assignable to $^2t_2g \longrightarrow {}^2Eg$ transition for O_h symmetry[19]. The electronic spectrum of complex V (III) exhibits band at 16100 Cm^{-1} with a shoulder at 20,400 Cm^{-1}. The low energy band has been assigned to $^2A_1g \longrightarrow {}^3A_2g$ while the high energy band may be due to $^2A_1g \longrightarrow {}^3T_2g(P)$ transition. These bands are characteristic of Octahedral geometry[20].

The electronic spectrum of Mn (III) complex showed an intense and sharp charge transfer band at 19800 Cm^{-1} and a spin allowed d-d transition band $^5Eg \longrightarrow {}^5T_2g$ at 18230 Cm^{-1}. This broad band occurring at lower frequency with increased intensity indicates the lowering of symmetry from Octahedral Configuration[21].

The electronic spectrum of the complex of Fe (III) exhibited these bands at 11300, 21600 and 27880 Cm^{-1} assignable to $^6A_1g \longrightarrow {}^4T_1g$, $^6A_1g \longrightarrow {}^4T_2g$ and $^6A_1g \longrightarrow {}^4Eg$ transitions respectively. These transitions are characteristics of Octahedral Fe (III) complexes [22].

The electronic spectrum of Co (III) complex displays bands at 15110, 21000 and 23300 Cm^{-1} assignable to $^3A_1g \longrightarrow {}^3T_2g$, $^1A_1g \longrightarrow {}^1T_1g$ and $^1A_1g \longrightarrow {}^1T_2g$ transitions respectively. These are similar to those reported for other six coordinated Co(III) complexes [23].

The electronic spectrum of complex of Ru (III) shows three bands at 13700, 17650 and 22500 Cm^{-1}. These bands are assigned to $^4T_g \longrightarrow {}^4T_{1g}$, $^2T_{2g} \longrightarrow {}^4T_{2g}$ and $^2T_{2g} \longrightarrow {}^2A_{2g}$, $^2T_{1g}$ respectively. These are similar to those reported for other Ru (III) octahedral complexes[24].

The electronic spectrum of Ru (II) complex shows a band at 22000 Cm^{-1}, this band has been assigned to charge transfer transition arising from excitation of an electron from metal t_{2g} level to unfilled molecular orbitals derived from π^* level of the ligands in accordance with the assignment made for similar octahedral Ru (II) complexes[25].

The electronic spectrum of oxovanadium (IV) complex exhibited three bands at 11870, 19425 and 26120 Cm^{-1} assignable to $dxy(b_2) \longrightarrow dxy$, $dyz(e^*$, $dxy(b_2) \longrightarrow dx^2\text{-}y^2(b_1)$ and $dxy(b_2) \longrightarrow dz^2\ (a_1)$ transitions respectively for octahedral stereochemistry[26].

I. R. spectra

The IR data of the spectra of Schiff base ligand and its complexes are presented in Table-2. The IR spectra of the complexes were compared with those of the free ligands in order to determine the coordination sites that may be involved in chelation. There were some guide peaks in the spectra of the ligands, which were helpful in achieving this goal. The position and/or the intensities of these peaks are expected to change upon chelation. New peaks are also guide peaks, as is water, in chelation. Upon comparison, it was determined that the□ ν(C=N) stretching vibration is found in the free ligand at 1615. This band was shifted to higher wavenumbers in the complex indicating the participation of the azomethine nitrogen in coordination (M-N)[27]. Medium to sharp bands, due to ν(C-O-C) stretching vibration of furan, appeared at 1229 cm^{-1} in the ligand[28]. This band shifted to 1274 cm^{-1} in metal complex[29]. These shifts refer to the coordination through a furan O atom.

New bands are found in the spectra of the complexes in the region 552 (furan O), which are assigned to ν (M-O) stretching vibration for metal complexes. The band at 420-453 metal complex have been assigned to ν(M-N) mode. Therefore, from the IR spectra, it is concluded that the ligand behaves as a neutral tetradentate ligand coordinated to the metal ions via azomethine N and furan O.

Table-2 IR data (4000-400 cm^{-1}) of Ligand and its metal complexes

Compound	ν (C=N)	ν (C-O-C)	ν (OH) (Water)	δ (H_2O) (Hydrated water)	ν (coordinated (M-O)	ν (M-N)
Ligand-L	1614sh	1229m	3349br	-	-	-
$[Fe(L)(H_2O)_2]Cl_3.3H_2O$	1628sh	1237m	3320sh	928m, 884m	552s	432m
$[Co(L)(H_2O)_2]Cl_3$	1626s	1232m	3401sh	931m, 884m	557s	452m
$[Ru)(H_2O)_2]Cl_3$	1611sh	1274m	3394br	928m, 883m	595m	420s
$[Cu(L)(H_2O)_2]Cl_3.2H_2O$	1617sh	1231sh	3381br	912m, 883m	657s	445m

Sh=sharp, m=medium, br=broad, s=small, w=weak.

Table-3 Proton NMR spectral data of Ligand

Compound	Chemical shift, (δ) ppm	Assignment
$C_{16}H_{12}N_2O_2$	8.102	(s, 2H, azomethine H)
	5.83 – 7.84	(m, 10H, 4ArH and 6 furan H)
	3.7	(br, 2H, H_2O)
	2.5	CH_3 of solvent)

Thermal analyses (TGA and Dr TG)

Thermogravimetric analyses (TGA and DrTA) of the Schiff base ligand and its chelates are used to : (i) get information about the thermal stability of these new complexes, (ii) decide whether the water molecules (if present) are inside or outside the inner coordination sphere of the central metal ion, and (iii) suggest a general scheme for thermal decomposition of these chelates. In the present investigation, heating rates were suitably controlled at 10^0C min^{-1} under nitrogen atmosphere, and the weight loss was measured from the ambient temperature up to ≅ 1000^0C. The data are provided in Table-4. The weight loss for each chelate was calculated within the corresponding temperature ranges. The TGA curve of Schiff base ligand exhibits a first estimated mass loss of 49.93% (calcd: 50.76%) at 30-400^0C, which may be attributed to the liberation of $C_8H_6O_2$ as gases. In the 3rd and 4th stages within the temperature range 400-900^0C, ligand loses the remaining part with an estimated mass loss of 50.07% (calcd: 49.24%) with a complete decomposition as CO, CO_2, NO, NO_2, etc. gases. The Fe (III) L. chelate shows 5 decomposition steps within the temperature range 30-1000^0C. The first 2 steps of decomposition within the temperature range 25-5000^0C correspond to the loss of water molecules of hydration and HCl, H_2, and O_2 gases, with a mass loss of 28.13% (calcd: 27.59%) for the Fe (III) L. chelate. The overall weight loss amounts to 86.01% (calcd: 84.52%) for the Fe(III) chelates with ligand (L). $[Cu(L)(H_2O)_2]Cl_2{\cdot}2H_2O$ chelate exhibit 1 to 4 decomposition steps. For the Cu(II)-L chelate, the first step is in the temperature range 30-120^0C (mass loss = 7.96%; calcd for $2H_2O$: 7.65%), which may account for the loss of water molecules of hydration. As shown in Table-4 the mass losses of the remaining decomposition steps amount to 22.85% (calcd: 22.74%) and correspond to the removal of HCl, H_2O, ½ O_2 L molecules, leaving CuO as a residue.

Table-4 Thermoanalytical results (TG, DrTG) of L and its metal complexes

Compound	TG range (^{0}C)	DrTG ^{max} (^{0}C)	n*	Mass loss Estim	Total mass loss (Calcd.) %	Assignment	Metallic residue
L	30-400	70, 267	2	49.93(50.76)	100.(100.0)	-Loss of $C_8H_6O_2$	-
	400-900	570, 758	2	50.07 (49.24)		-Loss of $C_8H_6N_2$	
	30-130	63	1	10.83 (10.46)		-Loss of $3H_2O$	
(1)	130-500	182, 320	2	28.13 (27.59)	86.01 (84.52)	-Loss of 3HCl,½ H_2 and O_2	½ Fe_2O_3
	500-1000	620, 880	2	47.05 (46.47)		-Loss of $C_8H_6N_2$	
	30-100	79	1	14.89 (14.34)		-Loss of $3H_2O$	
(2)	120-470	195, 280	2	22.83 (22.74)	83.90 (83.10)	-Loss of 2HCl,½ O_2 and H_2O	CuO
	470-850	740	1	53.11 (52.71)		-Loss of $C_{16}H_{12}N_2O$	
	30-100	79	2	7.58 (7.20)		-Loss of $3H_2O$	

N = number of decomposition steps (1) $[Fe(L)(H_2O)_2]Cl_3.3H_2O$ (2) $[Cu(L)(H_2O)_2]Cl_2.2H_2O$*

Calculation of activation thermodynamic parameters

The thermodynamic activation parameters of decomposition processes of dehydrated complexes, namely activation energy (E*), enthalpy (ΔH*), and Gibbs free energy changes of the decomposition (ΔG*), were evaluated graphically by employing the Coats-Redfern relation[30].

$$\log\left[\frac{\mathrm{Log)}W_f/W_f - W)\}}{T^2}\right] = \log\left[\left(\frac{AR}{E^*}\ \ 1 - \ \frac{2RT}{E^*}\right)\right] - \frac{E^*}{2.303RT} \qquad (1)$$

where W_f is the mass loss at the completion of the reaction, W is the mass loss up to temperature T, R is the gas constant, E^* is the activation energy in $kJ.mol^{-1}$, θ is the heating rate, and (1-(2RT/E*)) ≅ 1. A plot of the left-hand side of equation (1) against 1/T gave a slope from which E* was calculated and A (Arrhenius factor) was determined from the intercept. The entropy of activation (ΔS*), enthalpy of activation (ΔH*), and the free energy change of activation (ΔG*) were calculated using the following equations:

$$\Delta S^* = 2.30[\log(Ah/kT)]R \qquad (2)$$

$$\Delta H^* = E^* - RT \qquad (3)$$

$$\Delta G^* = \Delta H^* - T\Delta S^* \qquad (4)$$

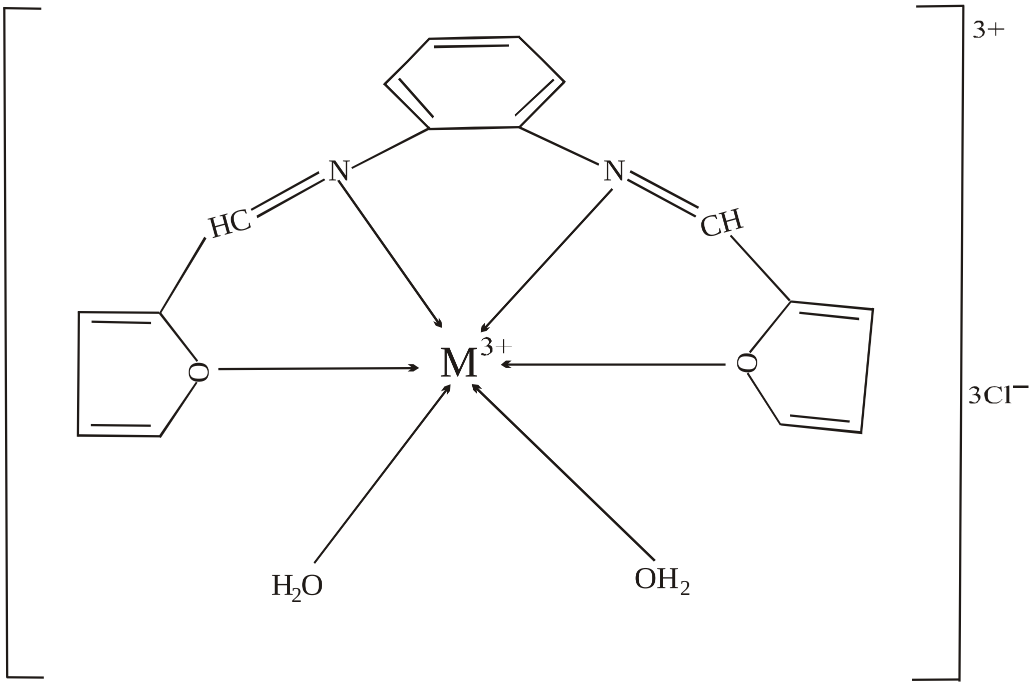
The data are summarized in Table-5. The activation energies of decomposition were in the range 55.42 – 350.6 kJ mol^{-1}. The high values of the activation energies reflect the thermal stability of the complexes. The entropy of activation had negative values in all the complexes, which indicates that the decomposition reactions proceed with a lower rate than the normal ones.

Table-5 Thermodynamic data of the thermal decomposition of metal complexes of L

Compound	Decomp.	E*	A	ΔS*	ΔH*	ΔG*
	Temp.(C)	kJ mol^{-1}	s^{-1}	KJ mol^{-1}	kJ mol^{-1}	kJ mol^{-1}
$[Fe(L)(H_2O)_2]Cl_3.3H_2O$	30-130	35.73	$1.23x10^{5}$	-108	73.48	72.65
	130-220	49.14	$4.05x10^{7}$	-136	61.49	96.2
	220-430	93.34	$4.51x10^{10}$	-224	98.39	48.68
	570-700	119.54	$6.23x10^{9}$	-104.6	101.6	66.5
	820-960	135	$5.01x10^{10}$	-202.6	96.43	88.08
$[Cu(L)(H_2O)_2]Cl_2.2H_2O$	30-120	80.8	$4.21x10^{10}$	-35.69	73.12	87.32
	120-210	145.7	$8.69x10^{14}$	-109.8	194.5	156
	210-450	223.7	$5.89x10^{14}$	-184.9	262.3	201.3
	500-800	267.2	$4.01x10^{10}$	-65.78	218.4	284.2

CONCLUSION

The structures of the complexes of Schiff bases L and complexes of Fe(III), Co (III), V (III), Cu (II), Ru (III), Mn (III) and Vo (IV) ions were confirmed by elemental analyses, IR, NMR, molar conductance, magnetic, solid reflectance, UV-Vis, mass and thermal analyses data. Therefore, from the IR spectra, it is concluded that L behaves as a neutral tetradentate ligand, coordinated to the metal ions via azomethine N and furan O. On the basis of above observations and magnetic and solid reflectance measurements, an octahedral geometry may be proposed for all the synthesized complexes.



Structure of Metal Complexes

M^{3+} = Ti (III), V (III), Mn (III), Fe (III), Co (III), Ru (III)

Figure-2

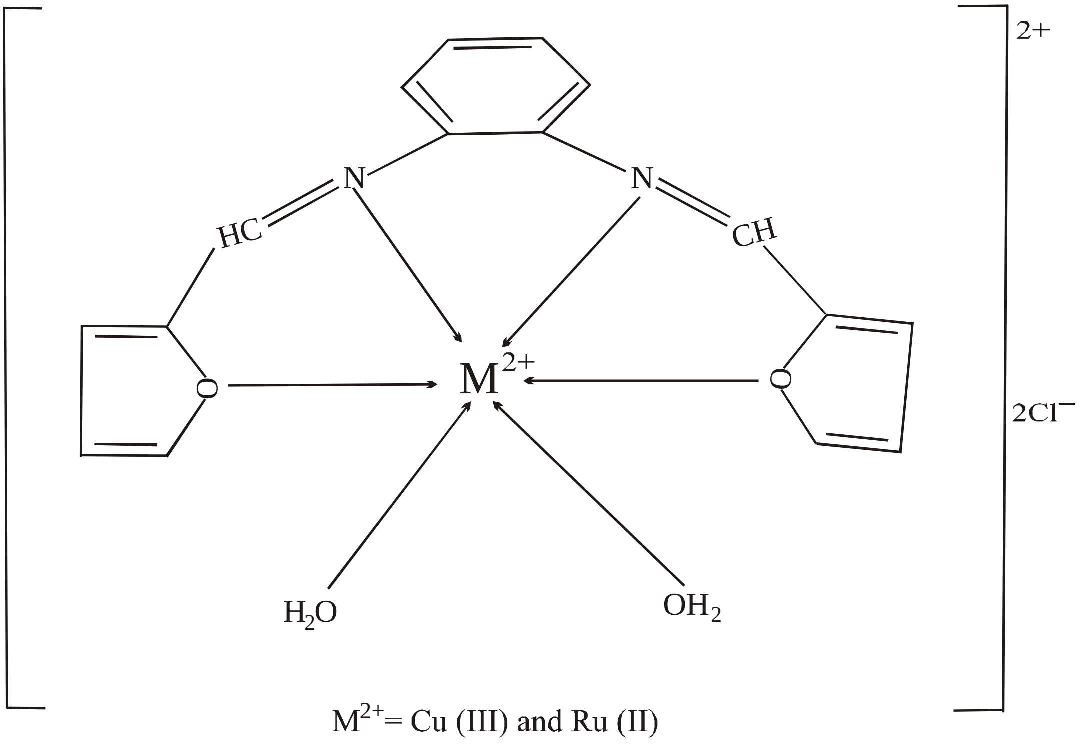


Figure-3

Biological activity

In testing the antibacterial activity of these compounds we used more than one test organism to increase the chance of detecting the antibiotic potential of the tested materials. The sensitivity of a microorganism to antibiotics and other antimicrobial agents was determined by the assay plates, which were incubated at 28^0C for 2 days (for yeasts) and at 37^0C for 1 day (for bacteria). All of the tested compounds showed a remarkable biological activity against different types of Gram-positive and Gram-negative bacteria. The data are listed in Table-6. Upon comparison of the biological activity of the Schiff base and its metal complexes with the standard (Traivid and Tavinic), it is seen that the biological activity of L are less than that of Tavinic, but higher than that of Traivid. For Schiff base (L) complexes, the biological activity of Fe (III), Co (III), Cu (II) and Ru (III) complexes is higher than that of the ligand and Traivid, while their activity is Fe(III), Co (III), Cu (II) and Ru (III) complexes is higher than that of the ligand and Traivid, while their activity is comparable with that of standard Travinic. The biological activity of the complexes follow the order Fe (III) = Co (III) = Cu (II) = Ru (III) > Mn (III) > Vo (IV) >Ni (II).

The importance of this lies in the fact that these complexes could reasonably be used for the treatment of some common diseases caused by E. coli, e.g., septicemia, gastroenteritis, urinary tract infections, and hospital-acquired infections[31].

Table-6 Biological activity of L and its metal complexes

Sample	*Staphylococcus pyogenes*			*Pseudomonas aeruginosa*			*Fungus (Candida) (Candida)*			*Escherichia*		
C, mg/L	5	2.5	1	5	2.5	1	5	2.5	1	5	2.5	1
L	++	++	+	++	+	+	+	-	-	++	+	-
$[Fe(L)(H_2O)_2]Cl_3.3H_2O$	+++	++	+	++	++	+	+	+	-	++	++	+
$[Co(L)]Cl_3.H_2O$	+++	++	+	++	++	+	+`	-	-	++	+	+
$[VO(L)(H_2O)_2]Cl_2.4H_2O$	++	+	-	++	+	+	+	-	-	+++	++	+
$[Cu(L)(H_2O)_2]Cl_2.2H_2O$	+++	++	++	+++	++	+	+	-	-	++	++	+
$[Mn(L)(H_2O)_2]\ Cl_3$	++	+	+	++	+	-	+	-	-	+++	++	+
$[Ru(L)(H_2O)_2]Cl_3$	+++	++	+	++	++	+	+	+	-	+++	++	+
Traivid	++	+	-	++	+	-	-	-	-	++	+	-
Tavinic	+++	++	+	+++	++	+	-	-	-	+++	++	+

The test was performed using the diffusion agar technique, Inhibition values = 0.1 – 0.5 cm beyond control = + Inhibition values = 0.6 – 1.0 cm beyond control = ++ Inhibition values = 1.1 – 1.5 cm beyond control = +++.

However, Fe (III), Cu (II), Ru (III), Co (III), VO (IV) and Cu (II) complexes of L ligand were specialized in inhibiting Gram positive bacterial strains (Staphylococcus pyogenes and Pseudomonas aeruginosa). The importance of this unique property of the investigated Schiff base complexes is that they could be administered safely for the treatment of infections caused by any of these particular strains. In addition, all metal complexes of L inhibit fungi at high concentration (5 mg/L), more so than the parent ligands and standards. Therefore, it is claimed here that such compounds might have a possible antitumor effect since Gram-negative bacteria are considered a quantitative microbiological method for testing beneficial and important drugs.

REFERENCES

[1] S. Patai, The chemistry of Carbon Nitrogen double bond, John Wiley & Sons Ltd., London, **1970.**
[2] Yamada S, *Coord. Chem. Rev*. **1999,** 537, 190-192.
[3] Singh S, Das S and Dhakarey R, *E-J Chem*., **2009,** 6(1), 99-105.
[4] Raman N, Syed Ali Fathima S and Dhaveethu Raja *J. Serbian Chem. Soc*. **2008,** 73 (II), 1063-1071.
[5] Joshi N, Balasubramanian G and Gharat LA, PCT Int. Appl. WO 2006051390; *Chem. Abstr*. **2006,** *144* , 488509**.**
[6] Tangallaplly RP, Lee R.E.B., Lenaerts AJM & Lee R.E. *Bioorg Med Chem Litt.*, **2006,** *16*, 2584.
[7] Srivastava V, Negi A.S., Kumar J.K., Faridi U, Sisodia B.S. Darokar M.P., Luqman S and Khaunja SPS, *Bioorg Med. Chem. Litt.* **2006,** *16***,** 911**.**
[8] Nagaraja GK, Kumaraswaney MN, Vaidya VP & Mahadevan KM, *ARKIVOC* (Geinesville, FL, united status), **2006,** *10*, 1**.**
[9] Shehata IA, *Saudi pharmaceutical journal* **2003**, *11*, 87; Chem. Abstr. **2004**, 140, 357263**.**
[10] Singh R.V., *Synth. React. Inorg. Met-org Chem*, **1986,** *16*, 21-27.
[11] Mohammad M.A., El-Enamy M.M., Basies E.L., *Egypt. J. Pharm. Sci.*, **1981,** *22*, 9-15.
[12] Scozzafava, A. Menabuoni, L. Mincione, G. Supuran, C.T., *Bioorg. Med. Chem. Litt.*, **2001,** *11*, 578-588.
[13] Sondhi S.M, Dinodia M. and Kumar A. *Bioorg. Med. Chem.* **2006,** 14, 4657.
[14] Sondhi S.M. Dinodia M. Singh J. and Rani R, *Current bioactive compounds*, **2007,** 3, 91.
[15] A.I. Vogel, "Practical organic chemistry including quantitative organic analysis", 3rd Ed., Longmans, London, **1956,** p. 854,
[16] D. feng and B. Wang, *Transition Met. Chem*. 18, 101 (13), (1993), M. Kumar, *Asian J. Chem*. **1994,** 6, 576-80.
[17] M.E. Ibrahim, A.A.H. Ali and F.M.M. Maher, *J. Chem. Techu. Biotechnol*. **1992,** 55, 217.
[18] N. Sari, S. Arsalan, E. Logoglu and Z. Sakiyan, *J of Sci*. **2003,** 16, 283.
[19] Mohammad Azim and Shamim Ahmad, *oriental J.C.*, **2011,** Vol. *27*, No. (2), 673-677.
[20] Rahul Kumar Rastogi, Poonam Garg and Shamim Ahmad, *Asian Journal of Chem*. **2009,** Vol. *21,* No.8 , 6144-6148.
[21] G.S. Bhadange, R.B. Mohod and A.S. Aswar, *Indian Journal of Chem*. **2001,** Vol. *40A*, PP 1110-1113.
[22] Patel M.M., Patel M.R., Patel M.N. and Patel R.P., *Indian J. Chem. Soc.* **1981,** *20A*, 6623.
[23] Chaudhary C.K., Chaudhary Ratan K and Mishra L.K. J. *Indian Chem. Soc.* **2003,** *80*, 693-695.
[24] Mahesh K. Singh, A.K. Singh, P.K. Gupta, Jaipal and L.K. Sharma I.*J. Chem*. **2002,** 41, 1385.
[25] R. Karvembu, K. Natrajan, *Polyhedron*, **2002,** 21, 219.
[26] Pancholi H.B., and Patel MM, *J. Polym. Moter*, **1996,** 13, 261.
[27] A.A. Soliman and W. Linert, *Thermochimica Acta*, **1999,** 333, 67-75.
[28] A. Kriza, M. Voiculescu and A. Nicolae, *Analele Universitatii Bucurestii*, Chemie, **2002,** 11, 197-201.
[29] J.K. Nag, D. Das, B.B. De and C. Sinha, *J. Indian Chem. Soc.* **1998,** 75, 496-498.
[30] A.M.S. El-sharief, M.S. Ammar, Y.A. Ammar and M.E. Zak, *Ind. J. Chem*. **1983,** 22B 700-704.
[31] A.M.S. El-Sharief, M.S. Ammar and Y.A. Mohammad, *Eggpt. J. Chem.* **1984,** 27, 535-546.